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Synthesis, characterization and antimicrobial studies on 13-membered-N₆-macrocyclic transition metal complexes containing trimethoprim

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Abstract Asymmetrical macrocyclic complexes of Mn^{II}, Co^{II}, Ni^{II}, Cu^{II} and Zn^{II} have been synthesized by the template process using bis(benzil)ethylenediamine as precursor. Bis(benzil)ethylenediamine reacts with transition metal chlorides and trimethoprim in a 1:1:1 molar ratio in methanol to give several solid metal complexes of the general composition [M(L)X₂] (M = Mn^{II}, Co^{II}, Ni^{II}, Cu^{II} and Zn^{II}, L = ligand and X = Cl⁻). They were characterized by physicochemical and spectroscopic techniques. Based on analytical, spectral and magnetic moments, all the complexes are identified as distorted octahedral structures. All the complexes are of the [M(L)X₂] type. The shifts of the ν(C=N) (azomethine) stretches have been monitored. To find out the donor sites of the ligands, the activity data show that the metal complexes are more potent than the parent ligand. The [M(L)X₂] complexes showed a broad spectrum of antimicrobial activity *in vitro* against both gram-positive and gram-negative human pathogenic bacterial isolates and the antimicrobial spectrum enhanced only with a combination of metal chlorides and trimethoprim complex. From the results it is imperative that the synthesized macrocyclic [M(L)X₂] complexes exhibit potent broad spectrum antibacterial activity.

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1. Introduction

The intense interest in synthetic macrocycles and their metal complexes depends on the fact that they mimic naturally occurring macrocyclic molecules in their structural and functional features due to rich chemical properties (Lindoy, 1989; Dietrich et al., 1993; Izatt et al., 1991). One of these properties



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is the enhanced thermochemical and kinetic stability of the complexes with regard to their dissociation which is due to lesser liability and larger association constants than the complexes with homologous open-chain chelating ligands. The application of polyazamacrocyclic precursors in the synthesis of transition metal macrocyclic complexes stems mainly from their use as models for protein-metal binding sites in biological systems (Kimura et al., 1985; Fenton and Okawa, 1993; Kimura, 1993; Karlin and Tyeklár, 1991) and as selective complexing agents for metal ions (Muller et al., 1983; Tsubuke et al., 1991; Beklemishev et al., 1994; Mohite et al., 1994).

In this paper, we report on the synthesis and characterization of new bivalent transition metal complexes from the reaction of bis(benzil)ethylenediamine with trimethoprim which show a marked effect of any build-up of ligand strain on the coordination geometry of the central metal. Due to emerging multidrug resistance to antibiotics like ampicillin (A), chloramphenicol (C), trimethoprim-sulfamethoxazole (Tm-Smz), ciprofloxacin (Cp), fluoroquinolones, (Fq) etc., treatment of infection becomes a problem during outbreaks of disease (Brusch, 2009; Dutta et al., 2008; Saha et al., 2006). Currently, the broad spectrum antibiotics were the preferred drugs of choice in many Asian countries, due to increased indiscriminate use of antibiotics for the treatment of humans and animals diseases, thereby developing multidrug resistance organisms (Mijovic et al., 2012), which pose serious threats in developing countries like India. With this view, the synthesized macrocyclic $[M(L)X_2]$ complexes were tested against human pathogenic bacterial isolates for evaluating their antimicrobial properties through disc diffusion and minimal inhibitory concentration (MIC) assay.

2. Experimental

2.1. Materials and methods

All the chemicals used for the preparation of the ligands were of BDH quality, AR grade. Molar conductance of the complexes was measured using a Toshniwals conductivity bridge using dip type platinized platinum electrode. Molecular weights were determined by the Rast camphor method. The magnetic moments were measured out by using the Gouy balance. Proton NMR spectra were recorded on a EM 300–30 MHz NMR spectrometer in DMSO. IR spectra (KBr) of the samples were recorded on a Shimadzu FTIR-8400s spectro-photometer. The electronic spectra (chloroform) were recorded on the Labomed UVD-3200 spectrophotometer (U.S.A.). The X-ray powder diffraction patterns were obtained with a SEIFERT model XRD 3000p powder X-ray diffractometer using Cu-K α radiation filtered by a nickel foil over the range of diffraction angle, $2\theta = 5\text{--}80^\circ$ in the complexes studied, where θ is the Bragg angle.

2.2. Synthesis of the complexes

To a methanolic solution of the precursor complex (Raghakrishna et al., 1996), bis(benzil)ethylenediamine, Trimethoprim (4.3 mM) and metal chlorides (4.3 mM) in methanol (25 ml) were added and the resulting mixture was heated under reflux for 5–6 h. The solution was cooled and transferred to an evaporating dish and set aside overnight at room temperature. The solid product that separated out was collected, washed with hot water, then with cold methanol and dried.

2.3. General properties

Mn(II), Co(II), Ni(II) and Cu(II) complexes are coloured but the zinc complexes are colourless (Table 1). All the complexes are soluble in dimethyl formamide (DMF), acetonitrile, DMSO and chloroform.

2.4. Bacterial strains

Five different bacterial strains used in this study were originally isolated from patient's samples. The bacterial isolates tested include gram positive species of *Staphylococcus aureus* and gram-negative species like *Escherichia coli*, *Shigella flexneri*, *Pseudomonas aeruginosa* and *Vibrio cholerae*. The bacterial strains were identified based on standard phenotypic, biochemical tests and grown overnight at 37 °C in nutrient broth. Cultures were maintained in nutrient agar (Hi-media, M002) slants at 4 °C. Subcultures were being made freshly before assaying.

2.5. Determination of antibacterial activity by disc diffusion assay

Antimicrobial activity of the synthesized macrocyclic $[M(L)X_2]$ complexes was carried out by the disc diffusion method. Two different concentrations of synthesized compounds (400–300 µg/ml) were prepared and their efficacy was assessed on Muller-Hinton agar (MHA) plates. Bacterial strains grown on nutrient broth medium at 37 °C for 24 h were adjusted to a final turbidity of 0.5 Mc Farland standard (10^8 CFU/ml) (Saad et al., 2011). A lawn culture was made with this bacterial load on the surface of MHA (Merck, U.S.A.). Aseptically, the synthesized $[M(L)X_2]$ complex was aseptically impregnated on a sterile 6 mm filter paper disc with 25 µl of each concentration and placed on bacterial seeded solid medium surface. All the plates were incubated at 37 °C for 24 h and tested in duplicate. The clear zone of inhibition around each disc was measured after incubation and recorded in millimetre. The standard antibiotics like, ampicillin (30 µg), nalidixic acid (30 µg),

Table 1 Physical, analytical data of the complexes.

Complex	Empirical formula	Colour	Mol. Wt.	M.P. (°C)	Yield (%)
$[Mn(L)Cl_2]$	$C_{44}H_{38}MnN_6O_3Cl_2$	Red	824	220	60
$[Co(L)Cl_2]$	$C_{44}H_{38}CoN_6O_3Cl_2$	Pink	828	242	55
$[Ni(L)Cl_2]$	$C_{44}H_{38}NiN_6O_3Cl_2$	Green	828	216	58
$[Cu(L)Cl_2]$	$C_{44}H_{38}CuN_6O_3Cl_2$	Pale blue	833	232	61
$[Zn(L)Cl_2]$	$C_{44}H_{38}ZnN_6O_3Cl_2$	Colourless	835	256	52

norfloxacin (10 µg), ciprofloxacin (5 µg), methicillin (5 µg), were included as control antibiotics and the zone of inhibition measured for these antibiotics were interpreted in accordance to manufacturers' standard (Hi-Media). The CLSI (NCCLS) standard method was employed to determine antimicrobial susceptibility to reflect the resistance rate of test pathogens for the tested compounds (Skov et al., 2000).

2.6. Determination of minimal inhibitory concentration (MIC) assay

The microdilution method was used to determine the lowest or minimum complex test material required for inhibiting the growth of the tested microorganism. A serial dilution of each macrocyclic [M(L)X₂] complex was prepared with dilution starting at 256–0.05 µg/ml using DMSO as solvent system (Irith et al., 2008). The MIC was performed in a 96 well assay plate filled with Muller-Hinton broth medium with different concentrations of the synthesized macrocyclic [M(L)X₂] complex with standard antibiotic as positive and DMSO solvent as negative controls. An equal volume of 100 µl of fresh bacterial broth suspension (10⁸ CFU/ml) was added to the wells without altering the dilution factor. The entire assay was prepared in triplicate in 96 well MIC plates and incubated at 37 °C for 24 h. The minimum inhibitory concentration (MIC) was determined by comparing the various concentrations of the synthesized macrocyclic [M(L)X₂] complex showing different inhibitory effect and selecting the lowest concentration that inhibits the growth of the tested microorganism (Prasannabala et al., 2012).

2.7. Determination of minimum bactericidal concentration (MBC)

To determine the Minimum Bactericidal Concentration (MBC), the lowest concentration of the synthesized macrocyclic [M(L)X₂] complex or standard antibiotic which destroys the tested bacterial strains in the MIC assay was cultured on freshly prepared MHA plates and incubated at 37 °C for 24 h. The highest diluted compound showing no growth on agar plates after incubation was considered as the MBC value.

3. Results and discussion

The elemental analysis is shown in Table 2. The molar conductance values obtained for these complexes at the concentration of 10⁻³ M are in the range of 10–50 Ω⁻¹ cm² mol⁻¹. These values are too low to account for any dissociation of the complexes in acetonitrile. Hence these complexes can be regarded as non-electrolytes. The ligand (L) is soluble in common organic solvents such as THF, DMSO and methanol. The octahedral metal complexes are highly soluble in DMSO, DMF and chloroform.

3.1. Conductance

The observed molar conductance values of the complex in acetonitrile are in the range of 10–50 Ω⁻¹ cm² mol⁻¹ at room temperature. Hence from conductivity measurements, it is concluded that the chloride ions are covalently bonded to metal ions, which indicates that they act as ligands and not as simple ions. Based on the metal-ligand ratio calculated by the analytical data and the nature of the electrolytes given by the conductance measurements, compositions were assigned for the prepared complexes. From the magnetic and conductometric analysis it is predicted that the complexes may have the following structures [Mn(L)Cl₂], [Co(L)Cl₂], [Ni(L)Cl₂], [Cu(L)Cl₂] and [Zn(L)Cl₂].

3.2. Magnetic moments

The magnetic moment values of Co(II), Ni(II) and Cu(II) complexes are shown in Table 2. The magnetic moment of Co (II) complex is in the range of 5.01 BM indicating that the Co(II) complexes are typically high spin complexes having an octahedral structure. The Ni(II) complexes exhibit the magnetic moment values in the range 3.01 BM, indicating octahedral coordination of the ligands around Ni(II) ion. The Cu(II) complexes exhibit magnetic moment in the range 2.01 BM suggesting a distorted octahedral nature. The absorption bands observed for the electronic spectra of the metal complexes also support the octahedral geometry.

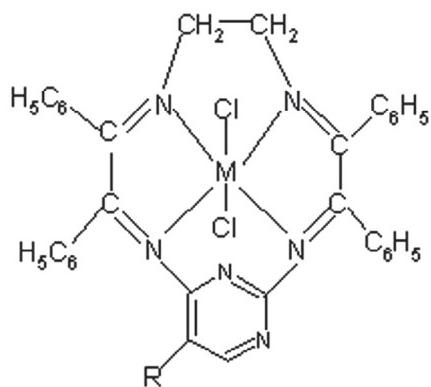
Table 2 Elemental analysis, magnetic moment, and molar conductance of the complexes.

Ligand/complex	Elemental analysis % found (calculated)				$\mu_{\text{eff.}}$ BM	Molar Conductance (Ω^{-1} cm ² mol ⁻¹)
	C	H	N	M		
C ₄₄ H ₃₈ MnN ₆ O ₃ Cl ₂	64.01 (64.08)	4.55 (4.61)	10.10 (10.19)	6.59 (6.66)	5.94	26.15
C ₄₄ H ₃₈ CoN ₆ O ₃ Cl ₂	63.69 (63.77)	4.52 (4.59)	10.09 (10.14)	7.07 (7.11)	5.01	31.14
C ₄₄ H ₃₈ NiN ₆ O ₃ Cl ₂	63.71 (63.77)	4.56 (4.59)	10.10 (10.14)	7.01 (7.09)	3.01	26.68
C ₄₄ H ₃₈ CuN ₆ O ₃ Cl ₂	63.31 (63.39)	4.52 (4.56)	10.00 (10.08)	7.57 (7.62)	2.01	28.21
C ₄₄ H ₃₈ ZnN ₆ O ₃ Cl ₂	63.19 (63.23)	4.51 (4.55)	10.03 (10.06)	7.82 (7.86)	Dia	30.5

3.3. Infrared spectral analysis

The IR spectra of Ligand (L) with its octahedral complexes have been studied in order to characterize their structures. The IR spectra of the free ligand and its metal complexes were carried out in the 4000–400 cm⁻¹ range. The infrared spectra of the ligand C₃₀H₂₄N₂O₂ and the metal complexes show the absence of uncondensed functional groups (NH₂ and C=O), stretching modes of the starting material and the appearance of bands characteristic of the imine group (Shakir et al., 1995). The bands characteristic of the benzil moiety appeared in all the complexes at 1500–1551 cm⁻¹ ($\nu_{\text{asym}}\text{C}_6\text{H}_5$) and

1316–1399 cm⁻¹ ($\nu_{\text{sym}}\text{C}_6\text{H}_5$). The major changes observed in the IR spectra of the macrocyclic complexes are the absence of stretching and deformation vibrations of NH₂ groups, indicating their deprotonation and the appearance of strong bands due to coordinated $\nu(\text{C}=\text{N})$ vibrations (Kumari et al., 1993), in the range 1635–1644 cm⁻¹. Strong and sharp bands for C–H stretching and bending vibrations appear in 2835–2837 and 1444–1453 cm⁻¹, respectively (Coltrain and Jackels, 1981). The presence of new bands in the spectra of the metal complexes in the far-IR region at 419–439 cm⁻¹ due to the $\nu(\text{M}=\text{N})$ vibrations supports the coordination of the imine nitrogen to the metal ion (Singh et al., 1989). Fig. 2 describes the FT-IR spectrum of the Co(C₄₄H₃₈N₆O₃)Cl₂ complex.



Where,

R = 3,4,5-trimethoxy benzyl

M = Mn^{II}, Co^{II}, Ni^{II}, Cu^{II} and Zn^{II}

Figure 1 Suggested structure of the complex.

3.4. Electronic spectra

The electronic spectra of bis(benzil)ethylenediamine and its derivatives have been recorded on a Lambda 35 spectrometer. The absorption maximum at 375 nm in the case of bis(benzil) ethylenediamine may be assigned to the $n \rightarrow \pi^*$ transition of the azomethine group. The shift of this band (15 nm) in the spectra of the complexes indicated the coordination of the nitrogen to the metal atom (Dixit and Tandon, 1990; Malik et al., 1983). In addition to this band, the spectra of the metal complexes exhibit bands around 270 and 335 nm due to $\pi \rightarrow \pi^*$ electronic transitions. However, the position of these bands remains almost unchanged on complexation. The UV-visible absorption spectral data of the Schiff base metal complexes are given in Table 3, which includes the absorption regions, band assignments and the proposed geometry of the complexes. These values are comparable with other reported values (Larabi et al., 2003).

3.4.1. Manganese(II) complexes

The Mn(II) complex shows a magnetic moment of 5.94 BM at room temperature corresponding to the five unpaired electrons. The diffuse reflectance spectra of Mn^{II} tetraaza

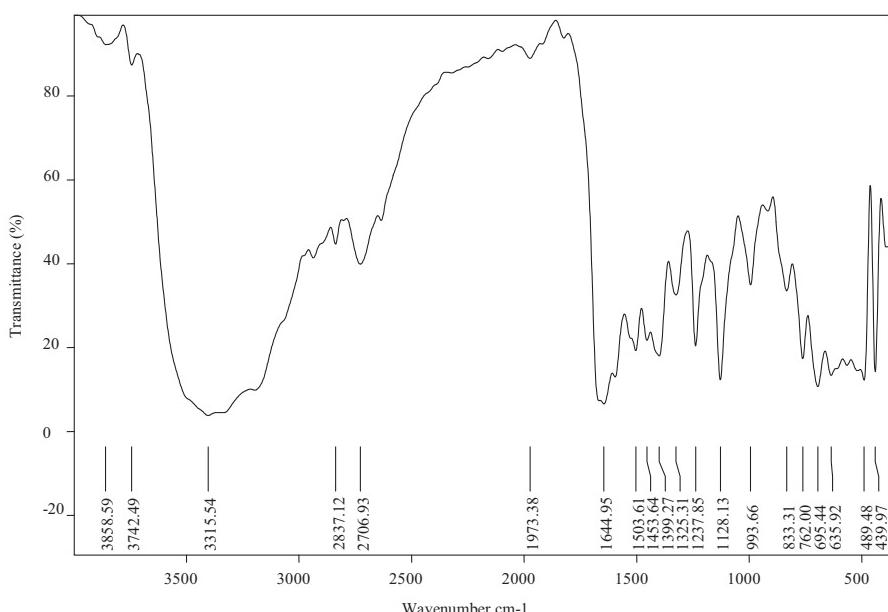


Figure 2 FT-IR spectrum of the Co (C₄₄H₃₈N₆O₃) Cl₂ complex.

Table 3 UV–Vis absorption data of the complexes.

Complex	λ_{max} (cm $^{-1}$)	Band assignments	Geometry
C ₄₄ H ₃₈ MnN ₆ O ₃ Cl ₂	18,811 31,296	$^6\text{A}_{1g}(\text{S}) \rightarrow ^4\text{T}_{1g}(\text{S})$ $^6\text{A}_{1g}(\text{S}) \rightarrow ^4\text{T}_{1g}(^4\text{P})$	Octahedral
C ₄₄ H ₃₈ CoN ₆ O ₃ Cl ₂	9302 14,669 17,886	$^4\text{T}_{1g}(\text{F}) \rightarrow ^4\text{T}_{2g}(\text{F})$ $^4\text{T}_{1g}(\text{F}) \rightarrow ^4\text{A}_{2g}(\text{F})$ $^4\text{T}_{1g}(\text{F}) \rightarrow ^4\text{T}_{1g}(\text{P})$	Octahedral
C ₄₄ H ₃₈ NiN ₆ O ₃ Cl ₂	10,394 15,869 28,253	$^3\text{A}_{2g} \rightarrow ^3\text{T}_{2g}$ $^3\text{A}_{2g} \rightarrow ^3\text{T}_{1g}(\text{F})$ $^3\text{A}_{2g} \rightarrow ^3\text{T}_{1g}(\text{P})$	Octahedral
C ₄₄ H ₃₈ CuN ₆ O ₃ Cl ₂	18,164	$^2\text{E}_g \rightarrow ^2\text{T}_{2g}$	Tetragonal

macrocyclic complexes exhibit bands in the region 14,280 cm $^{-1}$ which may be a charge transfer band. Bands in the region 18,811 and 31,296 cm $^{-1}$ may be assigned to $^6\text{A}_{1g}(\text{S}) \rightarrow ^4\text{T}_{1g}(\text{S})$ and $^6\text{A}_{1g}(\text{S}) \rightarrow ^4\text{T}_{1g}(^4\text{P})$ multiplicity forbidden transitions respectively (Satpathy et al., 1992). Thus the electronic spectral findings favour the geometry to be octahedral for Mn^{II} complexes.

3.4.2. Cobalt(II) complexes

The electronic spectrum of the cobalt complexes exhibited three bands at 9302, 14,669 and 17,886 cm $^{-1}$ that are assigned to the transitions $^4\text{T}_{1g}(\text{F}) \rightarrow ^4\text{T}_{2g}(\text{F})$, $^4\text{T}_{1g}(\text{F}) \rightarrow ^4\text{A}_{2g}(\text{F})$ and $^4\text{T}_{1g}(\text{F}) \rightarrow ^4\text{T}_{1g}(\text{P})$ respectively, suggesting octahedral geometry around the Co(II) ion. The magnetic moments of Co(II) complexes were found to be 5.01 BM also indicating octahedral geometry.

3.4.3. Nickel(II) complexes

The absorption spectra of Ni(II) complexes display three d - d transition bands at 10,394, 15,869 and 28,253 cm $^{-1}$ which correspond to $^3\text{A}_{2g} \rightarrow ^3\text{T}_{2g}$, $^3\text{A}_{2g} \rightarrow ^3\text{T}_{1g}(\text{F})$ and $^3\text{A}_{2g} \rightarrow ^3\text{T}_{1g}(\text{P})$ (Fig. 4). The magnetic moments of Ni (II) complexes were found to be 3.01 BM supporting the d^8 high spin distorted octahedral structure.

3.4.4. Copper(II) complexes

The magnetic moment values of Cu(II) complexes are 2.01 BM that fall within the normal range observed for distorted octahedral complexes. From the results, Cu (II) complexes show a single broad band in the range 18,164 cm $^{-1}$ due to transition between $^2\text{E}_g \rightarrow ^2\text{T}_{2g}$ suggesting tetragonal geometry. Tetragonal or square planar Cu (II) complexes are expected to give three bands. However, these three bands usually overlapped in tetragonal complexes, to give one broad absorption band. The electronic spectra and magnetic moment data for all Cu(II) complexes coupled with the analytical, conductance data suggest the tetragonal geometry for all the complexes.

3.5. EPR spectral studies

3.5.1. EPR spectrum of Cu(II) complex

The EPR spectra of the Cu(II) complexes were recorded in a polycrystalline sample as well as in solution at room temperature in different frequencies. The analysis of spectra gives the values of $g_{\parallel} = 2.1755$ and $g_{\perp} = 2.0568$. The observed ‘ g ’ values for the complexes are less than 2.3 in agreement with the covalent character of the metal ligand bond. The trend $g_{\parallel} > g_{\perp} > 2.0023$ observed for the complexes indicates that the unpaired electron is localized in $d_{x^2-y^2}$ orbital of the Cu(II) ion and spectral features are characteristic of axial symmetry. Tetragonal elongated structure is confirmed for the aforesaid complex.

3.5.2. ^1H NMR spectra

The bonding pattern in the resulting complexes has been further substantiated by the proton magnetic resonance spectra of the precursor and the metal complexes of the macrocycles. The ^1H NMR spectra of the complexes do not show any signal corresponding to primary amino protons. This suggested that the proposed macrocyclic skeleton has been formed. A singlet observed at δ 3.709 ppm in the complex may be assigned to nine methoxy protons in the 3,4,5-trimethoxy benzil group of Trimethoprim. Another singlet is observed at δ 3.608 ppm which corresponds to benzil protons. A Singlet peak is observed at δ 3.533 ppm for the methylene protons. The shift

Table 4 X-ray powder pattern reports of the [Cu(C₄₄H₃₈N₆O₃)Cl₂] complex.

Complexes	2θ values	Unit cell parameters	Density (gcc)	Possible geometry
[Cu(C ₄₄ H ₃₈ N ₆ O ₃)Cl ₂]	17.392 21.742 24.038 24.889 25.872 26.952 28.444 32.16 32.758 34.731 38.907 45.388 46.429 49.617 53.844	$a = 9.424 \text{ \AA}$ $b = 14.680 \text{ \AA}$ $c = 9.326 \text{ \AA}$ $\alpha = 90.000^\circ$ $\beta = 109.32^\circ$ $\gamma = 90.000^\circ$	2.02	Monoclinic

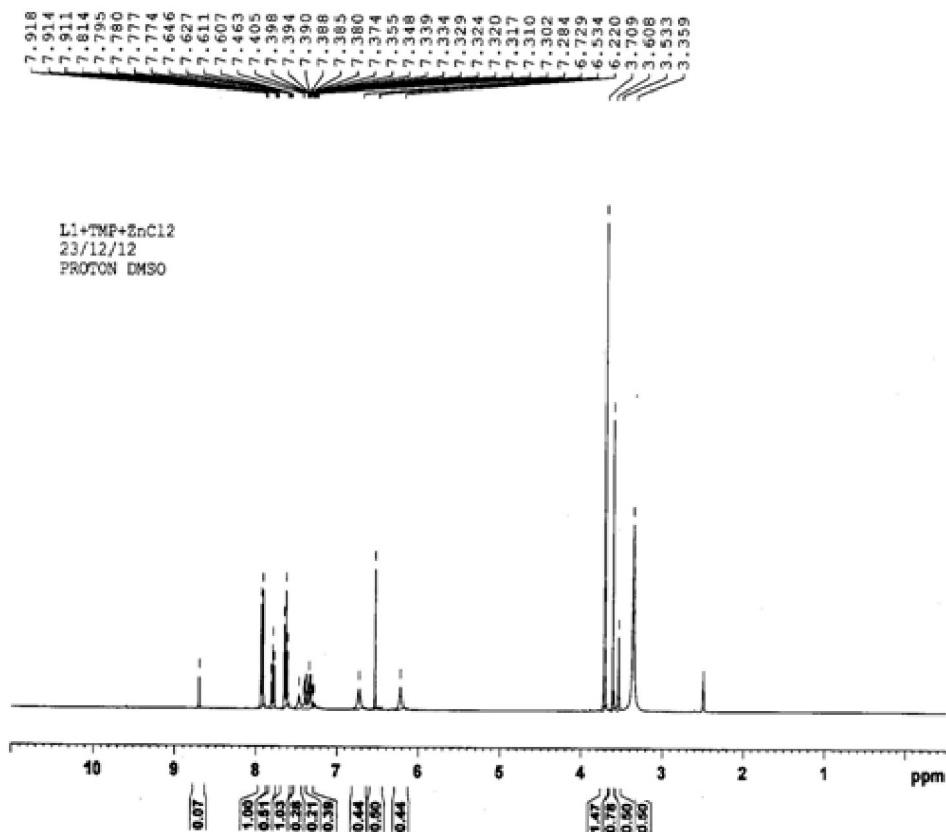


Figure 3 ^1H NMR spectrum of the $\text{Zn}(\text{C}_{44}\text{H}_{38}\text{N}_6\text{O}_3)\text{Cl}_2$ complex.

of the signals towards the lower field is an indication of the coordination of the macrocycles. The multiple of aromatic protons was observed at δ 7.3–7.9 ppm in the spectra of the precursor and the metal complexes of the macrocycles (Fig. 3).

3.6. Powdered XRD analysis

The diffractogram of the Cu^{II} complex consists of fifteen reflections with maxima at $2\theta = 24.889^\circ$ corresponding to the value of $d = 3.574 \text{ \AA}$ (Fig. 5) between 10° and 80° . The unit cell of Cu(II) complex yielded values of lattice constants: $a = 9.424 \text{ \AA}$, $b = 14.680 \text{ \AA}$, $c = 9.326 \text{ \AA}$ and a unit cell volume $V = 1290.199 \text{ \AA}^3$ (Table 4). In concurrence with these cell parameters, conditions such as $a \neq b \neq c$ and $\alpha = \gamma = 90^\circ \neq \beta$ required for a monoclinic sample are tested and found to be satisfactory. Hence, it can be concluded that the Cu(II) complex belongs to monoclinic crystal system.

The experimental density of the complex under study is determined by using the specific gravity method, which further enabled to calculate the volume of the unit cell. By using this value of density, molecular weight of the complexes, Avagadro's number and volume of unit cell, the number of molecules (n) per unit cell is calculated by using the equation, $\rho = nM/NV$. The experimental density values of the complex are determined using the specific gravity method and found to be 2.02 g cm^{-3} for this complex (Çakir and Biçer, 2004). By using this value of density, molecular weight of the complex, Avagadro's number and volume of unit cell, the number of molecules (n) per unit cell is calculated by using the equation, $\rho = nM/NV$ and was found to be 2 for this complex (Gale, 1996). With

this number, theoretical density is computed and found to be 2.04 g cm^{-3} for this complex. Comparison of experimental and theoretical density values shows good agreement within the limits of experimental error (Allan et al., 1979).

Thus from the above observations, the following structure is tentatively proposed for the macrocyclic complexes. The proposed structure is shown in Fig. 1.

3.7. Antimicrobial activity of the complex compound

Results of the antimicrobial sensitivity study for the macrocyclic $[\text{M(L)}\text{X}_2]$ complex by the disc diffusion test against all five human pathogenic bacterial strains are depicted in Table 5. Among the five $[\text{M(L)}\text{X}_2]$ complexes, the $[\text{Co(L)}\text{Cl}_2]$ complex exhibits highest bacteriostatic and bactericidal activity towards all pathogenic isolates other than *Vibrio* sp.

Among the tested bacterial strains, *E. coli*, *P. aeruginosa* and *V. cholerae* were susceptible to all tested standard antibiotics, while the gram-positive *S. aureus* and gram-negative *S. flexneri* bacterial strains showed resistance to methicillin and nalidixic acid standard antibiotics with a zone of inhibition of 9 and 12 mm respectively. The study of disc diffusion assay indicated that $[\text{Co(L)}\text{Cl}_2]$ is the most active complex with a zone of inhibition of 36 mm (400 $\mu\text{g}/\text{ml}$) against *S. flexneri* and 32 mm (400 $\mu\text{g}/\text{ml}$) against *S. aureus* [$\text{Zn(L)}\text{Cl}_2$] macrocyclic compound showed moderate antibacterial activity (Table 5). The macrocyclic $[\text{Mn(L)}\text{Cl}_2]$ complex did not show much effect against all bacterial isolates tested showing moderate activity only against *S. aureus*. Similarly $[\text{Cu(L)}\text{Cl}_2]$ showed low to moderate activity against *S. aureus* and *S. flexneri*. Of

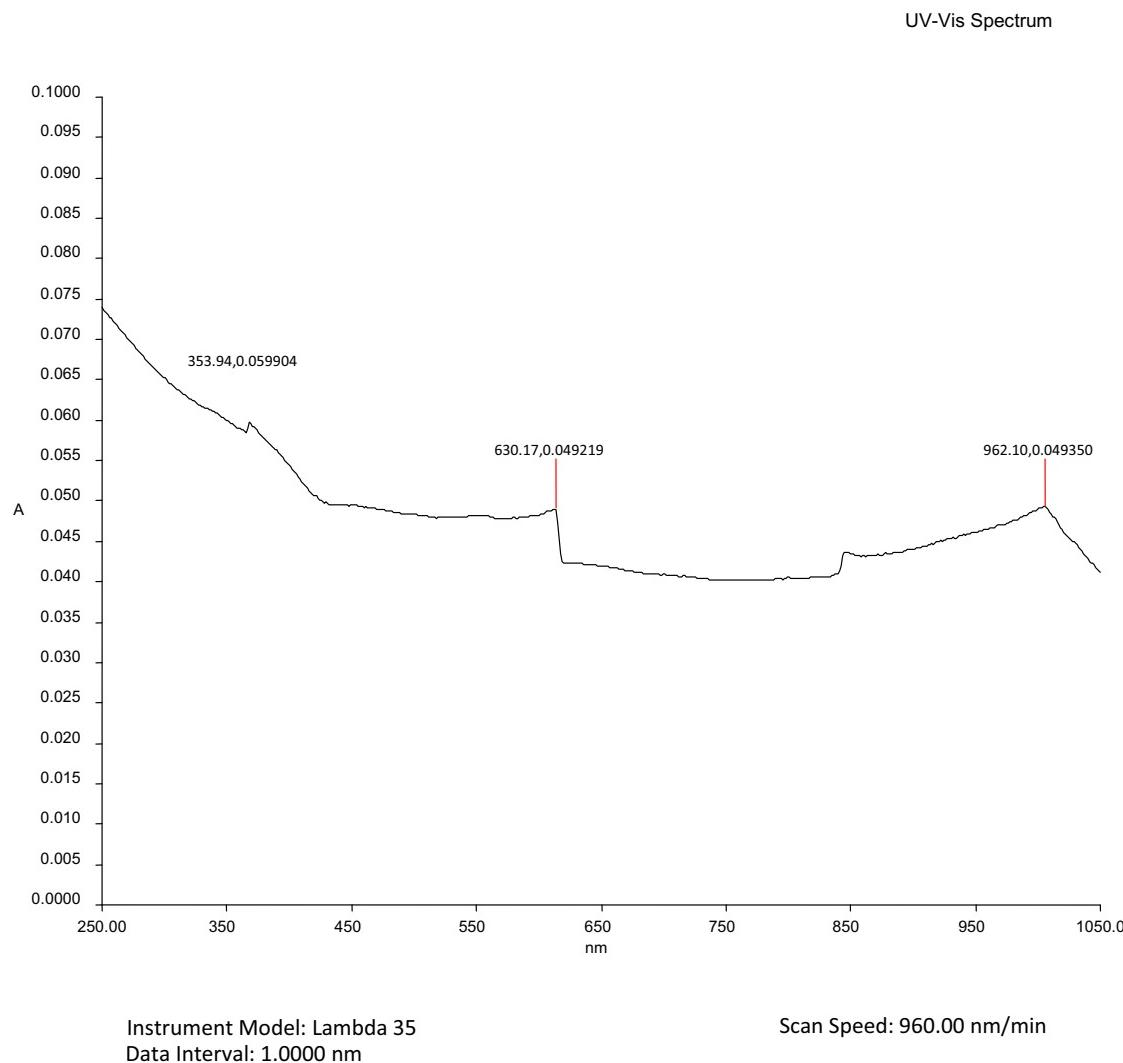


Figure 4 UV-visible spectrum of the $[\text{Ni}(\text{C}_{44}\text{H}_{38}\text{N}_6\text{O}_3)\text{Cl}_2]$ complex.

the four complexes tested by disc diffusion assay, $[\text{Co}(\text{L})\text{Cl}_2]$ complex expresses effective bacteriostatic activity against methicillin resistant *S. aureus* (MRSA) and nalidixic acid resistant *S. flexneri* pathogenic isolates. Antimicrobial resistance against the first line of drugs like nalidixic acid, trimethoprim and other antibiotics was commonly observed in many countries worldwide during *in vivo* analysis of pathogens (Rahman et al., 2007; Salam and Bennish, 1988). Similar work done by other researchers with a moderately different macrocyclic ligand and its metal complexes reported antimicrobial activities towards *S. aureus*, *E. coli* and *P. aeruginosa* bacterium which is in accordance with our macrocyclic complexes antibacterial potency (Ahmed et al., 2013). To know the bacteriostatic and bactericidal activities of the synthesized compounds, MIC assay was carried out. All four $[\text{M}(\text{L})\text{X}_2]$ complexes were used at a concentration from 256 to 0.5 $\mu\text{g}/\text{ml}$ against the human pathogenic bacterial isolates tested. The MIC results are shown in Table 6. Out of the five complexes tested $[\text{Co}(\text{L})\text{Cl}_2]$ showed better activity with an MIC value of 4, 16, 32 and 256 $\mu\text{g}/\text{ml}$ against *S. flexneri*, *S. aureus*, *E. coli* and *P. aeruginosa* respectively, where as $[\text{Zn}(\text{L})\text{Cl}_2]$, $[\text{Cu}(\text{L})\text{Cl}_2]$, and $[\text{Mn}(\text{L})\text{Cl}_2]$ showed an MIC value of 256 $\mu\text{g}/\text{ml}$

against the pathogenic bacteria tested. The $[\text{Mn}(\text{L})\text{Cl}_2]$ complex showed activity only against *S. aureus*, the other bacterial pathogens are resistant to it. None of the synthesized complexes showed activity against *V. cholerae*. It is interesting to note that $[\text{Co}(\text{L})\text{Cl}_2]$ showed efficient bactericidal activity at a lowest MIC value of 4 and 16 $\mu\text{g}/\text{ml}$ against *S. flexneri* and *S. aureus* respectively which showed resistance to the standard antibiotic tested (Table 6).

The trimethoprim antibiotic antimicrobial activity did not show much active bacteriostatic or bactericidal activity against the tested pathogenic bacteria, while the activity was highly influenced by forming the macrocyclic Schiff-Base complex. The trimethoprim is a bacteriostatic antibiotic compound mainly used against the treatment of urinary tract infection and inhibits the reduction of dihydrofolic acid (Rang et al., 1999). Thymine utilized in DNA formation was generated by folate metabolism. The trimethoprim prohibits the formation of tetrahydrofolate from dihydrofolate and therefore halts folate synthesis (Heaslet et al., 2009). The $[\text{Co}(\text{L})\text{Cl}_2]$ complex showed MIC activity at 4 $\mu\text{g}/\text{ml}$ against the antibiotic resistant *S. flexneri*, which among other isolates showed a high rate of resistance to standard nalidixic acid antibiotic at 32 $\mu\text{g}/\text{ml}$

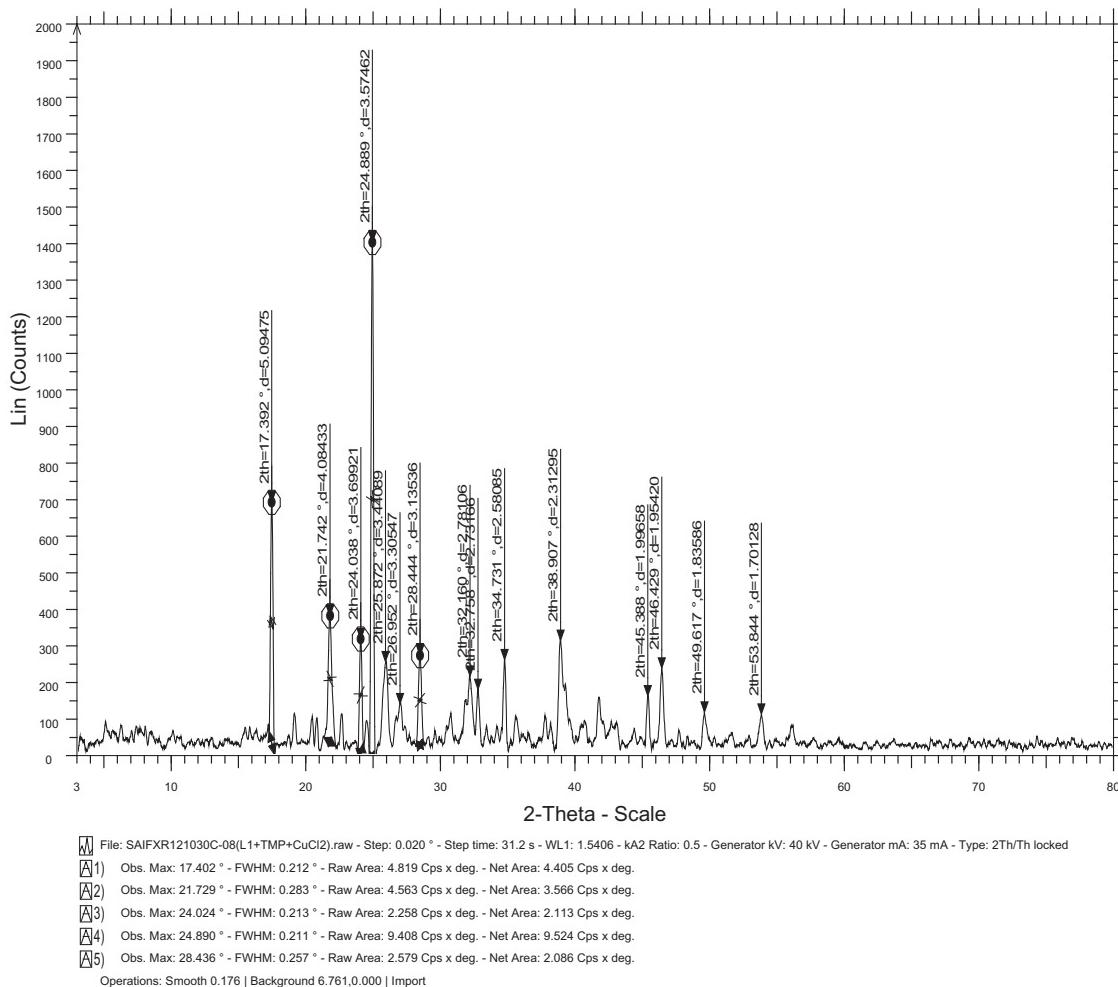


Figure 5 X-ray diffraction pattern of the $[\text{Cu}(\text{C}_{44}\text{H}_{38}\text{N}_6\text{O}_3)\text{Cl}_2]$ complex.

Table 5 The antibacterial activity by disc diffusion assay of the $[\text{M}(\text{L})\text{X}_2]$ complex against selected human pathogenic microorganisms.

Organism/[M(L)X ₂] complex	ZnCl ₂		MnCl ₂		CoCl ₂		NiCl ₂		CuCl ₂		AP	NX	CP	N	MT
	400 µg	300 µg	10 µg	10 µg	5 µg	30 µg	10 µg								
<i>Staphylococcus aureus</i>	24	13	24	10	32	24	NA	NA	23	15	29	23	30	31	9
<i>E. coli</i>	18	9	15	NA	20	19	10	NA	11	NA	30	32	34	27	—
<i>Shigella flexneri</i>	28	16	18	7	36	27	NA	NA	25	19	29	28	31	12	—
<i>Pseudomonas aeruginosa</i>	27	17	16	NA	29	16	9	NA	10	NA	32	18	34	28	—
<i>Vibrio cholerae</i>	10	NA	13	NA	19	7	NA	NA	12	NA	31	28	30	29	—

Bacteria tested zone of inhibition in mm.

AP, ampicillin; NX, norfloxacin; CP, ciprofloxacin; N, nalidixic acid; MT, methicillin; S* = sensitive; NA = no activity.

concentration (NCCL Breakpoint $\leq 8 \mu\text{g/ml}$). Similarly, methicillin resistance was demonstrated by tested *Staphylococcal aureus* isolate with an MIC value of $16 \mu\text{g/ml}$ (NCCL Breakpoint $\leq 8 \mu\text{g/ml}$).

The drug resistant bacterial strains possess resistance by the mechanism of enzymatic inactivation or target site modification with decreased intracellular drug accumulation (Schwart and Nobel, 1999). The rate of resistance to clinically isolated MRSA shows a lower level ($\leq 20\%$) of resistance to trimethoprim-sulfamethoxazole than other antibiotics (Kim et al.,

2004). The inhibition of nalidixic acid resistant *S. flexneri* and MRSA growth at lower concentrations indicated the higher antibacterial potency of our synthesized $[\text{Co}(\text{L})\text{Cl}_2]$ and $[\text{Zn}(\text{L})\text{Cl}_2]$ complex than known standard commercial antibiotics.

The $[\text{Ni}(\text{L})\text{Cl}_2]$ complex did respond less significantly to growth inhibition of *E. coli* and *P. aeruginosa* with a zone of inhibition of 10 mm (400 µg/ml) and 9 mm (400 µg/ml), while Fujiwara et al. (1990) express the neutral complexes of Ni(II) with Schiff bases exhibits the antimicrobial property for fungal

Table 6 The minimum inhibitory concentration assay of the macrocyclic [M(L)X₂] complex and standard antibiotics against human pathogenic isolates.

Organism/[M(L)X ₂] complex	MIC ($\mu\text{g/ml}$)									
	ZnCl ₂	MnCl ₂	CoCl ₂	CuCl ₂	Ni Cl ₂	AP	NX	CP	N	MT
<i>Staphylococcus aureus</i>	256	256	16	256	NA	S	S	S	S	64
<i>E. coli</i>	NA	NA	32	NA	NA	S	S	S	S	—
<i>Shigella flexneri</i>	256	NA	4	256	NA	S	S	S	32	—
<i>Pseudomonas aeruginosa</i>	256	NA	256	NA	NA	S	S	S	S	—
<i>Vibrio cholerae</i>	NA	NA	NA	NA	NA	S	S	S	S	—

Standard antibiotics: AP – ampicillin; NX – norfloxacin; CP – ciprofloxacin; N – nalidixic acid; MT – methicillin; S* = sensitive; NA = no activity.

species. The metal complexes possess greater antimicrobial properties against gram negative microorganisms compared to free ligands, which can be correlated to metal theory (Singh et al., 2004).

Our present study corroborates with some earlier reports on biological analysis of other metal complexes (Co(II), Cu(II), Ni(II), Mn(II) and Cr(III)) of macrocyclic Schiff-Base ligand, demonstrating both antibacterial and antifungal activities against *E. coli*, *S. aureus*, *Klebsiella pneumoniae*, *Mycobacterium smegmatis*, *P. aeruginosa*, *Enterococcus cloacae*, *Bacillus megaterium* and *Micrococcus leteu* (Prakash and Devjani, 2011). The Zn(II) complex showed a wide range of bactericidal activities against the gram positive and gram negative bacteria, with a potency well above the commercial antibiotics (Lei et al., 2009). Our study also authenticates the bacteriostatic and bactericidal nature of synthesized Co(II) and Zn(II) complexes. The results of the present investigation showed that the effective macrocyclic [M(L)X₂] complexes possess potent and broad-spectrum antimicrobial activity which could be utilized in the production of pharmaceutically important therapeutic antibiotics.

4. Conclusion

In this work, newly synthesized five different macrocyclic based metal complexes from the Schiff bases were investigated. All the five complexes, [Mn(L)Cl₂], [Co(L)Cl₂], [Ni(L)Cl₂], [Cu(L)Cl₂] and [Zn(L)Cl₂] were physicochemically characterized using elemental analysis, IR spectrum, UV, magnetic moments, ¹H NMR spectroscopies and powder XRD analysis. The elemental analyses correspond to the metal:ligand stoichiometry of 1:1. The electrical conductance value for the $\sim 10^{-3}$ M solutions of the complexes in acetonitrile, showed that all are non-electrolytic in nature. The infrared spectral data show that the ligand coordinates with the metal ion through the four imine nitrogen atoms of the macrocyclic ring. The magnetic moments of the complexes reveal that they are all of high spin type with paramagnetic (except Zn) nature. It suggests that Mn, Co, Ni, Cu and Zn are in the +2 oxidation state in their complexes. The electronic spectral data support that all the complexes were found to have octahedral/distorted octahedral geometry. The integral intensities of each signal in the ¹H NMR spectra of the macrocyclic complexes are found to agree with the number of different types of protons present. The X-ray diffraction analysis of Cu(II) complex showed that this complex belongs to the monoclinic crystal system. This is in concurrence with the cell parameters. The conditions such as $a \neq b \neq c$ and

$\alpha = \gamma = 90^\circ \neq \beta$ required for a monoclinic sample were tested and found to be satisfactory. The complex contains trimethoprim with two chloride ions and exhibits a distorted octahedral structure. All the macrocyclic complexes were evaluated for antimicrobial property against five human pathogenic bacteria namely *E. coli*, *P. aeruginosa*, *Vibrio cholerae*, methicillin resistant *S. aureus* and nalidixic resistant *S. flexneri*. Among the five macrocyclic Schiff base complexes, the [Co(L)Cl₂] and [Zn(L)Cl₂] complex exhibits potent antibacterial activity against drug resistant *S. flexneri* and *S. aureus* on comparing with standard drugs. Our data indicated that these complexes were effective to combat the growth of selective drug resistant pathogenic microorganisms.

Conflict of interest

The authors declare that they have no competing interest.

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